
Christensen PA, AbdHalim BMA, Mashhadani ZTAW, Martin PA. [A Direct Fourier Transform Infrared Spectroscopic Comparison of the Plasma- and Thermally-Driven Reaction of CO₂ at Macor](#). *Plasma Chemistry and Plasma Processing* 2018.

DOI link

<https://doi.org/10.1007/s11090-018-9874-6>

ePrints link

<http://eprint.ncl.ac.uk/245864>

Date deposited

06/02/2018

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A Direct Fourier Transform Infrared Spectroscopic Comparison of the Plasma- and Thermally-Driven Reaction of CO₂ at Macor

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Received: 8 November 2017 / Accepted: 14 January 2018
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Abstract This paper reports the appraisal of two in situ Fourier Transform InfraRed plasma cells with respect to the interrogation of the glow of a non-thermal plasma (using a transmission cell), and the non-thermal plasma/solid (i.e. dielectric/catalyst) interface (with a reflectance cell). The paper also reports, for the first time, a direct comparison of the IR spectroscopy of plasma- and thermally-driven chemistry. The system chosen for study was the reduction of CO₂ as there is a wealth of data in the literature for comparison. The catalyst was Macor, a ceramic material comprising primarily Al, Si and Mg oxides. In both the thermal and plasma experiments, rotationally-excited CO₂ (CO₂^{*}) was observed: in the plasma system, rotationally-excited CO (CO^{*}) was produced via the reduction of CO₂. Using the transmission cell, the conversion of CO₂ to CO was estimated and found to be up to 9% at energy efficiencies of ca. 1–2%, in line with the literature. No reaction of CO₂ was observed in the thermal system. The data obtained using the reflectance cell were similar to those obtained with the transmission cell, with the minor differences reflecting the longer residence time and higher specific input energy. Interestingly, two plasma-induced bands were observed in the reflectance experiments which increased in intensity with time and input power: these may be due transverse and longitudinal optical modes of SiO₂ and did not appear to participate in the observed chemistry.

Keywords Non-thermal plasma · In situ FTIR spectroscopy · Reflectance · Transmission · CO₂ reduction

Electronic supplementary material The online version of this article (<https://doi.org/10.1007/s11090-018-9874-6>) contains supplementary material, which is available to authorized users.

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Introduction

Non thermal plasma is produced by the partial ionization by electron impact of a gas using high AC or pulsed high voltages in the kV range. The temperature of the electrons in such plasmas are typically 10^4 – 10^5 K, whilst the heavier species such as the ions and gas molecules remain at temperatures around ambient. This generates radicals, ions and vibrationally and electronically excited species not normally present in thermal equilibrium at room temperature [1, 2]. In principle, almost 97% of the plasma discharge power of low temperature plasmas can be converted into vibrational excitation of the gas feed molecules [3, 4]. Adding a catalyst can speed up the chemical processes in a plasma and/or offer alternative mechanistic pathways with lower activation energies: thus, reactive species are produced and interact with catalyst at temperatures at which most thermally-driven process would be slow. Further, vibrationally excited molecules can have markedly different sticking probabilities and different chemistries on the surface of catalysts in contact with plasma [5] and it should be possible to enhance targeted reactions and suppress others by controlling the number density of electrons and electron and gas temperatures.

Inevitably, the first problem raised in any discussion of plasma-catalysis is low energy efficiency which is typically $\leq 10\%$ for the reactor designs commonly employed (usually tube reactors) [6]. In the longer term energy efficiency should improve once plasma catalysis is more fully understood. In the short term, however, this low efficiency need not be a major obstacle: for example, matching the production of electricity from renewable energy to demand is very difficult, and often the surplus is simply spilled: the storage of this surplus energy chemically is seen as a possible solution to this problem, and the plasma-driven reduction of CO_2 to CO and O_2 is perceived as one possible chemical process, particularly as some industrial plants produce almost pure CO_2 as waste gas [2, 7].

From the above discussion it is clear that, in principle, plasma-driven catalysed processes should open up wholly novel chemistries. However, progress has been slow, largely because catalyst selection appears to be based on those materials that are active for the analogous thermally-driven processes, and this may be due to the paucity of analytical information on non-thermal plasmas compared to, e.g. gas phase catalysis [1]. There are few analytical studies on the chemistry taking place actually in the glow of non-thermal plasmas, and these largely employ emission spectroscopy [8–10] or the detection of a single species [11]. There are a number of studies on the downstream analysis of the exhaust from NTPs, see for example [12], but actual studies of the plasma glow with, for example, IR spectroscopy [13] or of the catalyst surface in contact with a plasma is only a relatively recent phenomenon [14–18]. It is critical, if the potential of NTP is to be realized, that the identities of the reactive species on and near the catalysts surface are revealed [19]. In addition, there are no papers reporting direct comparison between thermal and plasma—driven catalyzed chemical conversions.

The aim of the work reported in this paper was to assess the efficacy of a combination of infrared reflectance and transmission spectroscopy to the study of non-thermal plasma and, in particular, the plasma/catalyst interface and to compare and contrast the data so obtained with the analogous thermally-driven process. The model system chosen was the reduction of CO_2 to CO and O_2 . Macor was chosen as the dielectric for its reasonable dielectric constant, ca 6, and its stability over the temperature range of interest, 25–600 °C [20] rather than for any possible catalytic activity.

Experimental

The Non-thermal Plasma Infrared Cells

The plasma transmission cell, see Fig. 1a, was designed and fabricated in-house and was made from a 15 cm long, cylindrical polytetrafluoroethylene (PTFE) tube with an outer diameter of 6.3 cm and an inner diameter of 2.5 cm. The high voltage electrodes were in the form of two plungers, sealing against the inner walls of the PTFE via rubber ‘O’ rings. The plungers were hollow down their axes (a 4 mm diameter hole) and the feed gas was delivered via one of these channels. The electrodes were covered with 5 mm thick Macor caps [Goodfellow Cambridge: quartz (35–50%), magnesium oxide (15–20%), aluminium oxide (15–20%) and fluoride (1–5%)] fitted snugly over each disc electrode: one cap had a hole aligned with the central, gas delivery channel and gas exhaust was via a nozzle on the top of the cell. Two infrared transparent CaF_2 windows (25 mm diameter, 3 mm thick, Crystran) were glued into the PTFE cell. The distance between the windows (path length) was 5.1 cm. The windows were positioned such that their centres aligned with the centre of the 5 mm gap between the electrodes. The plasma volume was 7.4 cm^3 and the residence time at a flow rate of $200 \text{ cm}^3 \text{ min}^{-1}$ was 2.2 s.

The electrodes were connected to a NeonPro lamp transformer, NP100000-30 (Hyrte, China) which provided an output of 10 kV at a constant frequency of 24 kHz. A voltage controller (Carroll & Meynell) was used to control the input power to the HV transformers. The input power to the plasma was monitored using a Gadget 13A power meter, N67FU (Maplin, UK); the input powers employed (20–28 W) were those obtained by subtracting the input power observed with plasma to the reading obtained with the system switched on but with no plasma initiated (4 W).

The reflectance NTP cell is shown in Fig. 1b: the IR beam passed through a 2 mm CaF_2 plate window and a 5 mm gap before being incident on the surface of a $3 \text{ cm} \times 3 \text{ cm} \times 0.5 \text{ cm}$ Macor plate covered by a Ti mesh (50% open area) electrode at an angle of 24° . The other electrode was a $2.3 \text{ cm} \times 2.2 \text{ cm} \times 0.007 \text{ cm}$ stainless steel foil mounted on an alumina plate behind the Macor. The same power supply etc. was employed as for the transmission cell. In contrast to the transmission cell, the plasma was in contact with a titanium mesh electrode as well as Macor. The plasma volume was 2.0 cm^3 and the residence time at a flow rate of $30 \text{ cm}^3 \text{ min}^{-1}$ (the cell could not sustain high flow rates) was 4.0 s.

For the transmission cell, the input powers of 20 and 28 W correspond to Specific Input Energies (SIEs) of 6.0 and 8.4 kJ dm^{-3} , respectively and defined as [21]:

$$\text{SIE (kJ dm}^{-3}\text{)} = \text{Discharge power (kJ s}^{-1}\text{)}/\text{Total gas flow rate (dm}^3\text{)} \quad (1)$$

Using the plasma reflectance cell, an input power of 24 W at a flow rate of $30 \text{ cm}^3 \text{ min}^{-1}$ corresponds to an SIE of 48.0 kJ dm^{-3} .

The design of the transmission cell did not allow for cooling of the electrodes apart from the $200 \text{ cm}^3 \text{ min}^{-1}$ flow of the feed gas: the reflectance cell had dedicated cooling channels for nitrogen gas, as shown in Fig. 1b.

The Plasma FTIR System

An Agilent FTS7000 FTIR spectrometer with a Deuterated TriGlycine Sulfate (DTGS) detector was employed. The IR beam was passed through the plasma via the two CaF_2

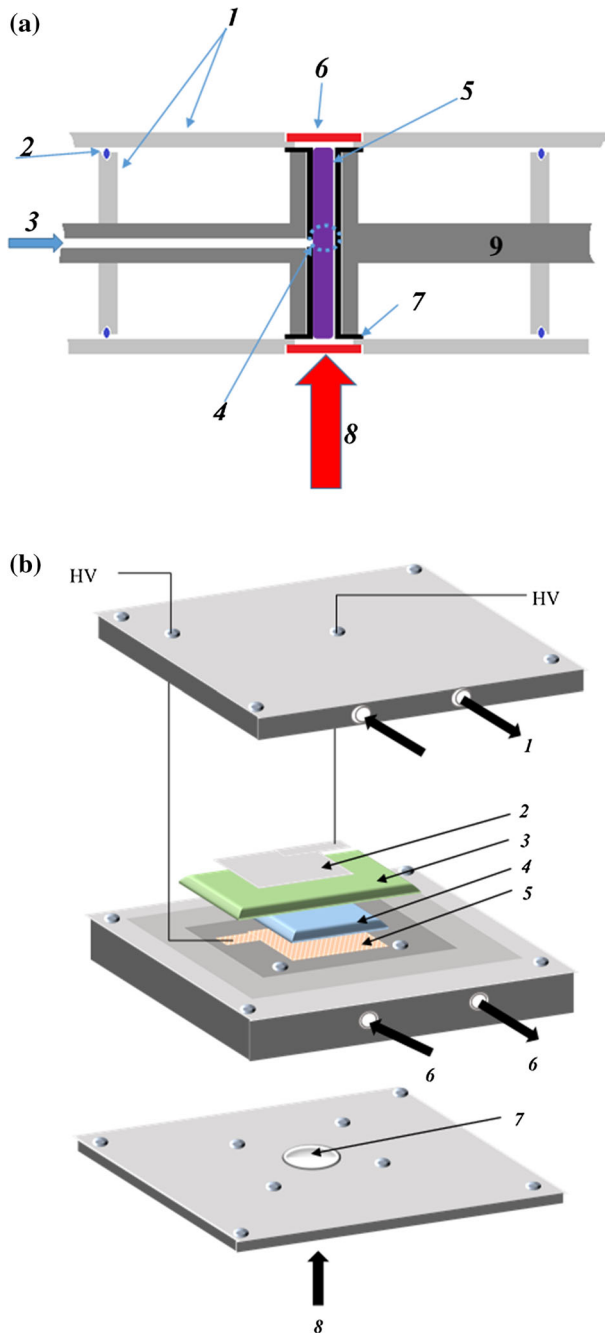


Fig. 1 The non-thermal plasma **a** transmission cell: 1 PTFE, 2 rubber “O” ring, 3 feed gas, 4 gas outlet, 5 plasma, 6 CaF_2 window, 7 Macor cap, 8 IR beam, 9 steel electrode and **b** reflectance cell: 1 cooling gas, 2 stainless steel shim, 3 Al_2O_3 , 4 Macor, 5 titanium mesh, 6 feed gas, 7 CaF_2 window, 8 IR beam

windows to the detector via an Amtir-1 filter, 25 mm × 2 mm (Spectra-Tech, USA) to remove visible light.

CO₂ (R744/100%) was supplied by BOC, UK and dinitrogen was obtained from a NITROSource generator. The gasses were mixed as appropriate and the composition controlled using flow meters. Table 1 summarizes the feed gas compositions employed in the work reported in this paper.

The cell was first flushed with N₂ at a flowrate of 200 cm³ min⁻¹ for 120 min, and a single beam spectrum taken (100 co-added and averaged scans at 8 cm⁻¹ resolution, 60 s per scan set), after which the CO₂ + N₂ at a nominal composition of 10% CO₂ + 90% N₂ was admitted to the cell at a total flow rate of 200 cm³ min⁻¹ and a second single beam spectrum collected, both without plasma. Sample spectra, S_S, were then taken every 2 min after the high voltage–power supply was switched on up to 20 min. By using the first single beam (of the nitrogen gas with no plasma) as the reference spectrum (S_R) all of the infrared active species present in the plasma were observed; using the second single beam as the reference spectrum resulted in difference spectra, showing only the changes incurred on initiating the plasma.

The Thermal FTIR System

Thermal in situ FTIR experiments were carried out using a Varian 670-IR spectrometer equipped with a ceramic air-cooled infrared source, a cooled DLaTGS detector and a Specac Environmental Chamber and diffuse reflectance unit [22]. A 12.5 mm diameter, 2 mm thick Macor disc was employed as the sample. The Specac reflectance accessory allows IR spectra to be collected under controlled atmosphere conditions from room temperature to 600 °C and pressures from vacuum to 34 at. The IR beam is incident on the sample in the cell at angles from 20° to 76° with respect to the horizontal plane via a ZnSe window. The reference spectrum (S_R, 8 cm⁻¹ resolution, 100 spectra per scan set, and 1 min per spectrum) was collected at 25 °C (see below) in the flowing N₂ + CO₂, after which the environmental chamber was isolated from the gas supply and sample spectra (S_S) were collected at 50 °C, and at 50 °C intervals thereafter up to 600 °C. The temperature was ramped at a rate of 5 °C min⁻¹.

Data Manipulation

The spectra obtained in the thermally-driven experiments, and plasma-driven experiments using the transmission cell are presented below as:

Table 1 The feed gas compositions employed in the plasma experiments reported in this paper

Input power (W)	Feed gas composition CO ₂ %, N ₂ % (Ar)
14	9.7/90.3
20	12.7/87.3
20	100/0
22	11.7/88.3
24	13.3/86.7
26	12.2/87.8
28	12.2/87.8
24*	31.1/68.9

*Reflectance cell

$$\text{Absorbance, } A = \log_{10}(S_R/S_S) \quad (2)$$

This results in difference spectra in which peaks pointing upwards (i.e. to + absorbance) represent a gain in absorbing species at S_S with respect to S_R , and peaks pointing down (to – absorbance) represent the loss of absorbing species. There was significant reflection from the window of the reflectance cell, and hence the single beam spectrum of the CaF_2 window was subtracted from single beam spectra S_S and S_R prior to the data manipulation in Eq. (2).

The concentrations of the various species observed in the plasma experiments were calculated using the Beer–Lambert law:

$$A = \varepsilon cL \quad (3)$$

where ε is the molar decadic extinction coefficient ($\text{M}^{-1}\text{cm}^{-1}$), c = concentration (M) and L = optical path length (cm).

Extinction Coefficients

The integrated extinction coefficient of $7.0 \times 10^5 \text{ cm mol}^{-1}$ reported by Bolis and co-workers [23] was employed for the CO absorption between 2002 and 2225 cm^{-1} : our methodology was checked by determining the extinction coefficient of CO using a 1 cm pathlength transmission cell with 100% CO at 24 °C, and the value so obtained compared to that of Bolis et al. and found to be within 10%. The same cell was employed to determine the integrated extinction coefficient of the combination bands of CO_2 (see below) between 3491 and 3769 cm^{-1} , and this was found to be $6.5 \times 10^5 \text{ cm mol}^{-1}$. Using the thermal FTIR system and a static atmosphere of 3.7% CO_2 + 96.3% N_2 , the integrated extinction coefficient of the CO_2 combination bands was found to remain constant over the temperature range from 25 to 150 °C of interest in this work.

Energy and Conversion Efficiency

In the plasma experiments, the conversion of CO_2 to rotationally-excited CO is:

$$\% \text{CO}^* = 100 \times [\text{CO}^*]_{\text{glow}} / [\text{CO}_2]_{\text{feed}} \quad (4a)$$

where $[\text{CO}_2]_{\text{feed}}$ is the concentration of CO_2 in the feed to the plasma cell and $[\text{CO}^*]_{\text{glow}}$ is the concentration of CO in the plasma glow. The corresponding CO selectivity, $\% \text{S}_{\text{CO}^*}$ is:

$$\% \text{S}_{\text{CO}^*} = 100 \times [\text{CO}^*]_{\text{glow}} / ([\text{CO}_2]_{\text{feed}} - [\text{CO}_2]_{\text{glow}}) \quad (4b)$$

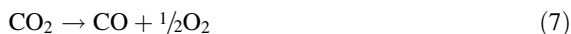
The carbon balance, $\% \text{C}$, is:

$$\% \text{C} = 100 \times ([\text{CO}^*]_{\text{glow}} + [\text{CO}_2]_{\text{glow}}) / [\text{CO}_2]_{\text{feed}} \quad (5)$$

The energy efficiency for the production of CO is:

$$\eta_{\text{CO}}/\% = 100 \Delta H / E_{p,I} \quad (6)$$

where ΔH is the enthalpy ($279.8 \text{ kJ mol}^{-1}$) of the process [7]:



at 300 K. $E_{p,\text{CO}}$ is the energy expended in producing 1 mol of CO and assuming that the IR spectra record an average picture over the 1 min data collection time.

All the calculations presented below are based on the integrated areas of the CO_2 and CO_2^* combination bands and the P and R branches of the CO^* absorption. These are all subject to inevitable experimental error, particularly with respect to determining baselines: this is particularly relevant to the determination of the integrated absorptions of the CO^* features as these are adjacent to the very strong CO_2^* fundamental. For the experiments employing the transmission cell, the estimated error in the baseline calculations of the CO^* bands, i.e. between employing a linear versus curved baseline between the extremes of the spectral ranges employed, was ca. 10%. The experimental error associated with the plasma reflectance cell was higher as the optical pathlength could not be measured directly (as was the case with the transmission cell) but was determined from calculations based on the optical configuration and the integrated intensity of the CO_2 combination bands of 100% CO_2 . On this basis, the pathlength of the reflectance cell was ca. five times lower than that of the transmission cell (5.1 cm eff. ca. 1 cm) making baseline correction even less accurate due to the weaker CO absorptions.

Results and Discussion

The Plasma-Driven Reaction of CO_2 at Macor

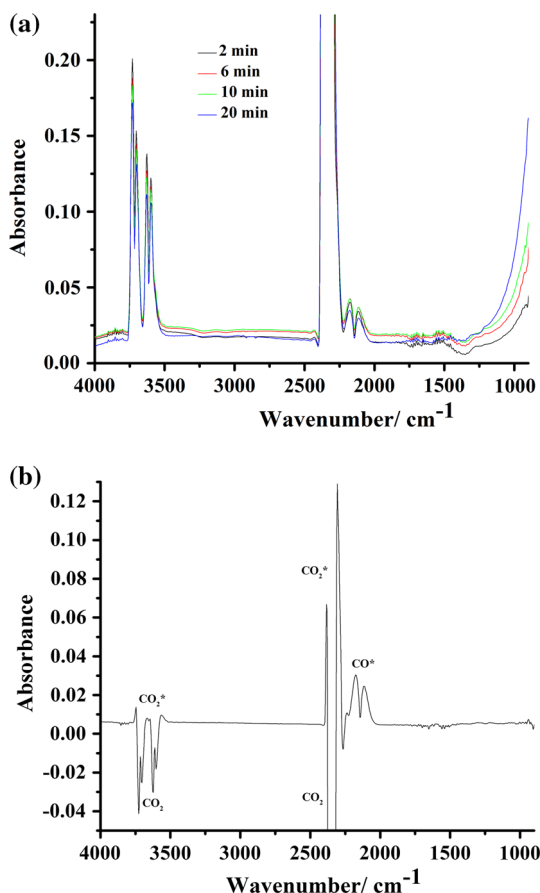
Plasma Transmission Cell

Considering first the data obtained using $\text{CO}_2 + \text{N}_2$, Fig. S1 shows a single beam spectrum of the plasma transmission cell at 25 °C with flowing 12.2% $\text{CO}_2 + 87.8\%$ N_2 and no plasma. The P and R asymmetric stretch (ν_3) bands of CO_2 at 2340 and 2360 cm^{-1} [24–26] are clearly saturated. The small feature around 2260 cm^{-1} may be attributed to the $^{13}\text{C}^{16}\text{O}_2$ fundamental [26]. The features between 3500 and 4000 cm^{-1} are attributable to CO_2 combination bands [27].

Figure 2a shows selected spectra collected as a function of time at 28 W input power employing the spectrum collected using nitrogen gas as the reference. The experiment in the figure was repeated at input powers of 20, 22, 24 and 26 W, with no additional bands being observed.

The saturating gain feature in Fig. 2a between 2250 and 2400 cm^{-1} is in the region of the P and R bands of the CO_2 asymmetric stretch. However, Fig. S2(a) shows the CO_2 ν_3 region of the spectra collected after 20 min in Fig. 2a, and the spectrum collected in the same experiment before the plasma was initiated: as can be seen, the spectra collected with plasma are somewhat broader than that collected without. Fig. S2(b) shows the analogous spectrum of the CO_2 combination band region after 20 min operation and that without plasma: although the intensities of the combination bands decreased on application of the plasma, it is clear that the bands are broadened. Given that the molecules present in the non-thermal plasma are at thermal equilibrium but the electrons are at a much higher temperature, then the spectra in Fig. 2a, S2(a) and S2(b) suggest that any CO_2 present in the plasma is rotationally excited [13], and that this species is responsible for the combination and asymmetric stretch features in Fig. 2a. Figure 2b shows the spectrum collected after 20 min in Fig. 2a using as the reference the single beam spectrum collected of

Fig. 2 **a** In situ FTIR spectra (8 cm^{-1} resolution, 100 co-added and averaged scans, 60 s per scan set) collected at the times shown on the figure at an input power of 28 W. The gas composition was 12.2% CO_2 + 87.8% N_2 and the reference spectrum was that taken with no plasma and a N_2 gas feed. **b** The spectrum collected after 20 min in **a** using the spectrum collected with no plasma and the 12.2% CO_2 + 87.8% N_2 feed gas as reference

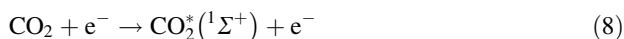


the CO_2/N_2 feed with no plasma: the CO^* bands are very clear, as are the structures between 2200 and 2400 cm^{-1} and 3500 and 3800 cm^{-1} due to the loss of ground state CO_2 bands superimposed upon the CO_2^* gain features. The features between 2200 and 2400 cm^{-1} in Fig. 2b were modelled using Spectralcalc and Origin and the difference spectrum obtained by subtracting the spectrum of CO_2 at room temperature from that of the excited CO_2 at 396 K are shown in Fig. S2(c): the figure supports the assignment of the gain features in Fig. 2a to higher temperature, rotationally excited CO_2 (CO_2^*). This assignment was further supported by the behaviour of the CO absorption in the plasma experiments and the thermal experiment, see below.

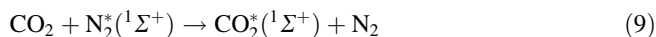
From Fig. 2a and S2(a) it is apparent that there are clear gain features due to the P and R bands of CO near 2116 and 2176 cm^{-1} [25, 28] and these are due to the higher temperature, rotationally excited gas, as may be expected from the discussion above. Thus, Fig. S3 shows the spectra collected after 2 and 20 min in the experiment in Fig. 2a, using the spectrum collected with the same gas feed but without plasma as the reference (see below) with a spectrum of CO gas in a 1 cm pathlength transmission cell at room temperature. The CO and 2 min spectra were scaled down to allow comparison. As may be seen from the figure, the CO bands in the spectra collected during the plasma experiment are somewhat broader than those of the room temperature CO, and the P and R maxima

move apart as would be expected for CO if the molecule was rotationally excited (CO*): for example, Bauerecker et al. [29] observed the rotation-vibration bandwidth of $^{13}\text{C}^{16}\text{O}$ to be reduced by about 40% on cooling the gas from 300 to 45 K, without any change in band centre. The resolution of the spectra obtained in the experiments reported in this paper (8 cm^{-1}) was chosen as the highest possible allowing for a reasonable data collection time. However, temperature calculations based on the maxima of the P and R branches of CO [30] gave wide ranges when the resolution was taken into account: e.g. allowing $\pm 4\text{ cm}^{-1}$ for P and R branches nominally at 2172 and 2118 cm^{-1} resulted in temperatures between 210 and 382 K. Hence, in order to allow comparison within the data sets, ambient temperature was employed.

The vibrationally excited form of the ground electronic state of CO_2 , $\text{CO}_2^*(^1\Sigma^+)$, is generally accepted as being the intermediate in the plasma-induced reduction of CO_2 to CO and O_2 for plasmas with lower mean electron energies than those in dielectric barrier discharges [3]. This is clearly not the same species as the rotationally-excited CO_2^* observed in Fig. 2. $\text{CO}_2^*(^1\Sigma^+)$ may be formed directly through collisions with electrons in the plasma [31, 32]:



or via collisional energy transfer from vibrationally excited species such as $\text{N}_2^*(^1\Sigma^+)$ [31]:



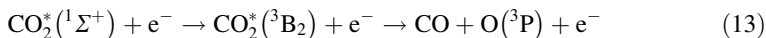
whilst $\text{CO}_2^*(^1\Sigma^+)$ decay can take place through relaxation to CO_2 and photon emission or dissociation into CO and atomic oxygen [26, 27]:



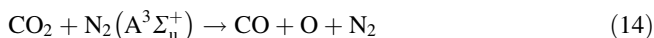
The dissociation of CO_2 via reaction (10) can take place by two possible routes [3, 33]: (1) excitation to vibrational levels of the ground excited state of CO_2 which exceed the energy of the dissociation threshold: this results in ground state CO and singlet oxygen:



(2) excitation to the first excited triplet state of CO_2 via intersystem crossing resulting in ground state CO and ground state atomic oxygen.



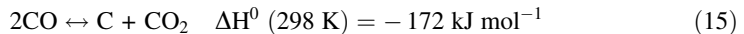
The latter is favoured due to the non-adiabatic intersystem crossing and the reduction in the energy required for dissociation [34, 35]. In CO_2 feeds diluted with nitrogen, Snoeckx et al. [6] postulate that the dissociation of CO_2 depends upon the partial pressure of N_2 : thus, at low nitrogen pressure, the main reaction is the direct electron excitation of CO_2 and subsequent reduction to CO and O_2 , as discussed above. In contrast, at high nitrogen partial pressure it is:



In other words, as the partial pressure of nitrogen increases, the electron energy is primarily employed to excite the nitrogen molecules to the metastable $\text{A}^3\Sigma_u^+$ state rather

than for direct electron excitation of CO₂, which may be the reason vibrationally excited CO₂ was not detected in our experiments.

The integrated absorption intensities of the CO₂ combination bands and the CO fundamental were employed to gain some quantitative insight into the chemistry of the plasma glow, and the results are summarized in Table 2. In general, the concentrations of both CO* and CO₂* declined with time at all input powers, leading to a loss in carbon inventory, as is shown in Table 2 and by the typical data (from the experiment in Fig. 2a) in Fig. 3 in which the integrated intensities of the CO* and CO₂* bands were normalized to their maximum values and plotted with respect to time: the raw data are shown in Fig. S4. The broadening of the CO₂ and CO features suggests the decline in the concentration of CO* is accompanied by a concomitant increase in the temperature of the gases in the plasma, and the two appear to be linked. The chemical simplicity of the plasma system and lack of any additional product absorptions suggests that there is a process, or processes, taking place leading to carbon-containing products that are infrared inactive or present at steady-state concentrations below the detection limit of the FTIR system. One example of the former is the Boudard reaction [36], which is commonly observed in plasma systems involving CO/CO₂:



No carbon deposits were observed on the walls or windows of the cell, but these could have been swept away by the feed gas. As was stated above, the plasma temperature increased with operational time and input power: due to the exothermicity of reaction (15) and its negative reaction entropy, the equilibrium constant for the process decreases with temperature, and hence less loss of CO due to the Boudard reaction would be expected with time, in contrast to the observed results.

As can be seen from Table 2, the conversion of CO₂ to CO after 2 min was ca. 10–11% irrespective of the input power, declining with time, as expected on the basis of Fig. 3 and S4. The former observation was unexpected as an increase in conversion with increasing plasma power would be qualitatively expected on the basis of an increasing frequency of electron/gas collisions due to an increase in the discharge current ($I = P/V_0$), as shown by Manley's equation [37]:

$$P = 4fC_dV_o(V - V_oC_g/C) \quad (16)$$

where P is the discharge power (W), V_o is the discharge onset voltage (kV), V is the discharge voltage (kV), C_g is the capacitance of the background gas within the discharge gap (F), C is the total capacitance of the dielectric and gas (F), ϵ_r is the dielectric constant, ϵ_o is vacuum permittivity ($8.854 \times 10^{-12} \text{ F m}^{-1}$), A and d are the dielectric surface area (m²) and thickness (m), respectively.

The IR plasma cell was in no way optimised for energy efficient operation, but a comparison of the energy efficiency for, for example, CO production with the literature values would at the least give confidence that the plasma system under study was not atypical, and this was confirmed. Thus Table 2 shows the values of η_{CO} observed during the various experiments. As may be seen from the table, the energy efficiency of CO production decreased with increasing operational time and increasing input power: e.g. at 20 W the η_{CO} declined from 2.4% after 2 min to 1.9% after 20 min, with the analogous values of 1.8–1.1% at 28 W. Overall, the efficiency for CO production were broadly comparable to those reported in the literature, i.e. ca. 1–2% [6, 7, 38, 39].

Table 2 Summary of the data obtained in the experiments typified by Fig. 1a

Input power (W)	Time (Min)	[CO ₂] _{feed} (10 ⁻³ M)	[CO ₂ [*]] _{glow} (10 ⁻³ M)	[CO [*]] _{glow} (10 ⁻⁴ M)	%CO [*]	η _{CO} (%)	S _{CO*} (%)	%C
20	2	5.2	4.7	5.1	9.9	2.4	102.0	100.0
	6		4.6	4.9	9.4	2.3	80.6	97.7
	10		4.5	4.6	8.8	2.1	66.9	95.6
	20		4.4	4.1	7.9	1.9	54.6	93.4
22	2	4.8	4.3	4.3	10.2	2.1	96.5	99.6
	6		4.1	4.1	9.3	1.9	70.2	96.0
	10		4.1	4.1	8.7	1.8	59.4	94.1
	20		4.0	4.0	7.8	1.6	50.3	92.3
24	2	5.4	4.9	5.1	9.4	2.0	89.7	98.9
	6		4.7	4.5	8.3	1.8	60.4	94.5
	10		4.6	4.2	7.6	1.6	51.1	92.7
	20		4.6	3.7	6.7	1.4	42.1	90.7
26	2	5.0	4.4	5.3	10.7	1.9	92.2	99.2
	6		4.2	4.5	9.1	1.6	62.2	94.5
	10		4.2	4.1	8.2	1.5	49.8	91.8
	20		4.1	3.5	7.0	1.2	39.7	89.4
28	2	5.0	4.3	5.3	10.7	1.8	75.9	96.6
	6		4.1	4.5	9.1	1.5	50.5	91.1
	10		4.0	4.0	8.1	1.3	42.3	88.9
	20		4.0	3.4	6.8	1.1	33.7	86.7
20 (100% CO ₂) [≠]	2	40.8	—	14.1	3.5	6.6	—	—
	6		—	13.0	3.2	6.0	—	—
	10		—	11.8	2.9	5.5	—	—
	20		—	9.8	2.4	4.6	—	—
24 (100% CO ₂) [≠]	2	40.8	—	—	—	—	—	—
	6		—	—	—	—	—	—
	10		—	—	—	—	—	—
	20		—	—	—	—	—	—
14 (Ar)	2	4.4	3.9	4.6	10.3	3.1	90.4	98.9
	6		3.8	4.2	9.5	2.8	68.3	95.6
	10		3.8	4.0	8.9	2.6	57.3	93.4
	20		3.7	3.4	7.7	2.3	43.9	90.2

[≠]CO₂ and CO₂^{*} absorptions were saturated, rendering calculation of their concentrations etc. not feasible

Fig. 3 Plots of the integrated absorbances of the CO* and CO₂* combination bands in Fig. 2a as a function of time. The plots were normalised to their maximum values

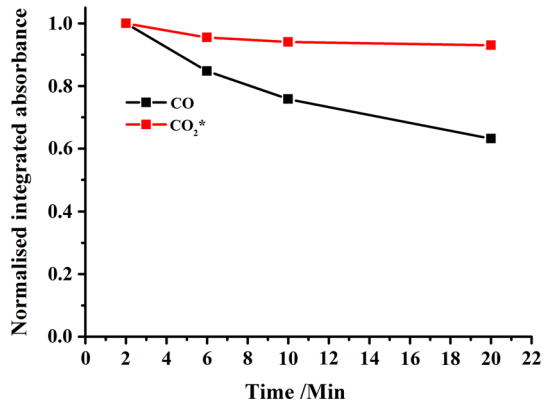


Table 2 also summarizes the data obtained using 100% CO₂ and CO₂ + Ar. With respect to the former, it can be seen that, whilst the increase in CO₂ results in an increase in the absolute amount of CO produced and the energy efficiency as would be expected, the fraction of CO₂ converted to CO decreases, from ca. 10% after 2 min to ca. 6–7%. These data support the positive role of N₂ in terms of its direct involvement in the reduction of CO₂, see (8)–(14) as well as the data obtained using Ar, see Table 2, where comparable conversion and energy efficiencies to those obtained using CO₂ + N₂ were obtained but at significantly lower input energy and ca. 20% lower CO₂ in the feed gas. Further, adding N₂ or Ar is generally reported as rendering plasmas easier to ignite and having significant effect upon electron energy distribution functions [40].

Plasma Reflectance Cell

Figure 4a shows selected sample spectra collected at an input power of 24 W and a nitrogen flow rate of 30 cm³ min⁻¹ using the plasma reflectance cell. The spectrum collected immediately before the plasma was initiated was employed as the reference and sample spectra were collected every 2–20 min after the plasma was turned on.

As may be seen from the figure, the spectra are featureless apart from two strong bands at 1210 (broad) and 1150 cm⁻¹, that increase in intensity steadily with time. Fig. S5 shows plots of the integrated absorptions of the two bands obtained from analogous experiments to that in Fig. 4a as a function of input power from the spectra taken after 20 min operation: clearly, the features also increase in intensity with input power.

Figure 4b shows the single beam spectra corresponding to the absorbance spectra in Fig. 4a along with the reference spectrum collected immediately before the plasma was initiated. As can be seen from the figure, the 1150 and 1210 cm⁻¹ features were present before the plasma was initiated and simply increase in intensity as the plasma is turned on and with time thereafter: this suggests that the features have their origin in the oxides on the surface of the Macor and/or Ti mesh. Interestingly, however, these features were not observed in the reflectance spectra of unused Macor + Ti mesh, suggesting they appeared in the presence of plasma in a previous experiment to that in Fig. 4b and had not completely relaxed by the time the latter commenced.

Figure 5 shows a repeat of the experiment in Fig. 4, using a feed gas composition of 49% CO₂ + 51% N₂. In addition to the features due to CO* and CO₂*, the bands at 1150 and 1210 cm⁻¹ are also present, and again grow with operational time.

Fig. 4 a In situ FTIR spectra (8 cm^{-1} resolution, 100 co-added and averaged scans, 60 s per scanset) collected at the times shown on the figure at an input power of 24 W using the reflectance cell and nitrogen gas as the feed gas. The spectrum collected immediately before the plasma was initiated was employed as the reference. **b** The single beam spectra from the experiment in **a**

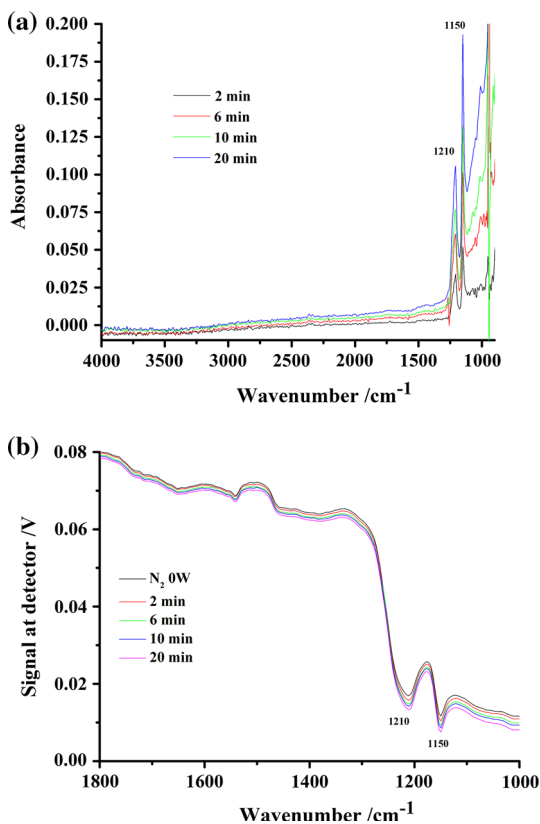


Fig. 5 In situ FTIR spectra (8 cm^{-1} resolution, 100 co-added and averaged scans, 60 s per scanset) collected at the times shown on the figure at an input power of 24 W using the reflectance cell and 49% CO_2 + 51% N_2 as the feed gas. The spectrum collected immediately before the plasma was initiated was employed as the reference

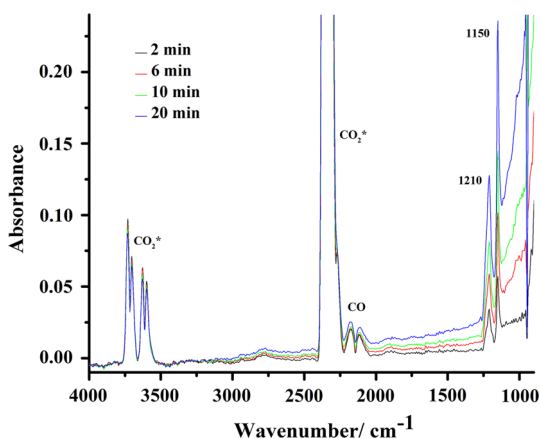


Fig. S6 shows plots of the normalised integrated absorptions of these features, and that of the ν_3 CO_2 asymmetric stretch, as a function of time after the plasma was switched off and the cell flushed with nitrogen. The absorptions were normalised to their values after 2 min flushing, to aid comparison. The absorptions of the two features do not track the CO_2 absorption, and the latter declines more rapidly: this suggests that the 1150 and 1210 cm^{-1}

bands are not due to a gas-phase species (as, perhaps, may be expected from the fact that they were not observed in the transmission cell experiments discussed above, although the presence of the Ti mesh is an additional factor in the reflectance experiments). The relatively slow ‘relaxation’ of these bands argues against their being due to some form of electric field enhancement of the Ti–O bonds on the mesh or the metal–oxygen bonds of the Macor, as such would be expected to relax as soon as the field was removed. In addition, the features also appeared in experiments where N_2 was replaced by argon; see, for example, Fig. S7 which shows an analogous experiment to that shown in Fig. 5, except using a feed gas composition of 31.0% CO_2 + 69.0% Ar and an input power of 14 W (when argon was employed in the feed, it was found that plasma was initiated and sustained at lower input power than when using N_2 , and the maximum power that could be employed was also lower); the 1150 and 1210 cm^{-1} features are clearly visible, growing with time. These features may be due to a transverse optical (TO) mode at 1210 cm^{-1} and a longitudinal optical (LO) mode at 1150 cm^{-1} of amorphous SiO_2 . These have previously been observed in the growth of thin films of SiO_2 [41] and were observed on the high wavenumber side of the large SiO_2 transverse optical peak due to the asymmetric stretch at 1070 cm^{-1} . They are thought to be due to disorder-induced mode coupling in amorphous SiO_2 with the exact positions varying dependent on conditions. In Lange’s work [41] they were shown to be stronger with ion bombardment. Whilst very interesting, it is not clear that the species responsible for the 1150 and 1210 cm^{-1} features participate in the reduction of CO_2 and hence are not relevant to the aim of this paper. Further work is in progress to elucidate this chemistry, and will be reported in a future publication.

Given the experimental uncertainties associated with the reflectance cell, and the fact that it is still under development, it is not possible to draw any meaningful, quantitative data from the spectra in Fig. 5: suffice to state that the production of CO_2^* fell from 92% of the CO_2 in the feed gas after 2 min operation to 89% after 20 min. However, the data in Fig. S4 strongly suggests that this decrease in the formation of CO_2^* was not reflected in an increase in the conversion to CO^* .

The absence of any features due to the oxides of nitrogen in the various spectra discussed above is worthy of note, as such species have been observed in CO_2 and N_2 plasmas [6, 42] and the specific energy densities employed by these authors were comparable to those employed in the current work. The absence of such species may be due to catalytic activity of the Macor and/or differences in construction of the reactors leading to different electric fields and electron energies.

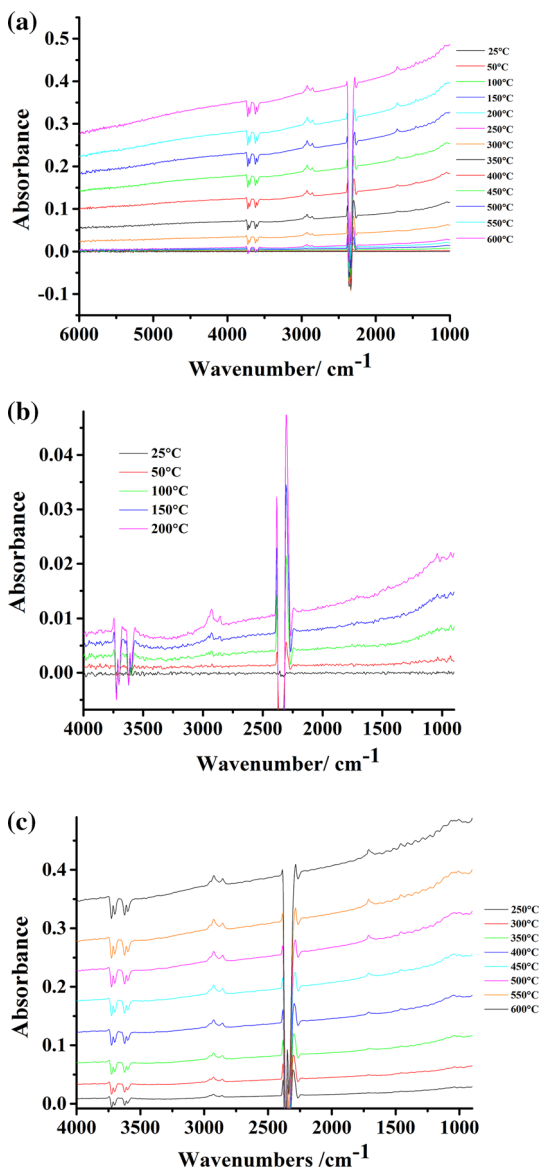
The Thermally Driven Reaction of CO_2 at Macor

The uncatalysed thermolysis of CO_2 into CO and O_2 takes place at negligible rates at temperatures < 1500 °C but the process is catalysed by metal oxides such as ZrO_2 , CeO_2 and mixed Ce/Zr oxides such that thermolysis occurs at temperatures > 1200 °C [43].

Fig. S8 shows spectra collected from the Macor disc under nitrogen up to 600 °C, showing only the spectra taken every 100 °C, for clarity. As can be seen from the figure, the spectra are dominated by a structured, broad loss between 3000 and 1000 cm^{-1} , the broad gain of an O–H stretch with a maximum around 3360 cm^{-1} , and sharp loss and gain features near 3700 cm^{-1} which may be attributed to isolated O–H stretches [44].

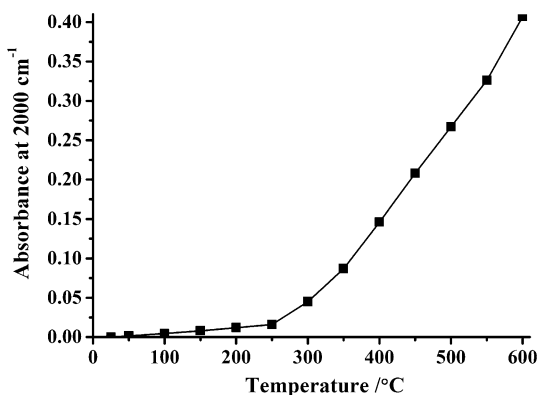
Figure 6a–c show the reflectance spectra collected in 23% CO_2 + 77% N_2 as a function of temperature up to 600 °C: Fig. 6a shows all the spectra, Fig. 6b the spectra up to 200 °C and Fig. 6c the spectra from 250 to 600 °C. For comparison, Fig. S9 shows the spectrum

Fig. 6 **a** In situ FTIR spectra (8 cm^{-1} resolution, 100 co-added and averaged scans, 60 s per scanset) collected from a 12.5 mm diameter Macor disc in a static atmosphere of 23% CO_2 + 77% N_2 at the temperatures shown on the figure. The temperature was ramped at $5\text{ }^\circ\text{C min}^{-1}$. The spectra in **a** up to **b** 200 $^\circ\text{C}$ and **c** 250 to 600 $^\circ\text{C}$



collected at 600 $^\circ\text{C}$ in Fig. S8 along with spectra collected at the same temperature in analogous experiments using 36% CH_4 + 64% N_2 and 21% CO_2 + 43% CH_4 + 36% N_2 , and Fig. 7 shows a plot of the absorbance at 2000 cm^{-1} in Fig. 6a versus temperature. It is clear from Figs. 6a and 5 that a broad, intense background feature extends across the spectrum and increases towards longer wavelength. This is likely to be due to radiative emission from the Macor disc which increases with temperature. However, it depends strongly on the composition of the gas phase above the Macor: in N_2 it is absent, it is intense and of equal intensity in 23% CO_2/N_2 and 21% CO_2 + 43% CH_4/N_2 , and present

Fig. 7 A plot of the absorbance at 2000 cm^{-1} of the spectra in Fig. 6a as a function of temperature



but of lower intensity in 36% CH_4/N_2 . Whilst extremely interesting, a discussion of these data are outwith the remit of this paper and will be presented in a subsequent submission.

Figure 6a–c are dominated by the very intense CO_2 loss features at 2360 and 2340 cm^{-1} , and the $\text{CO}_2/\text{CO}_2^*$ combination bands between 3500 and 4000 cm^{-1} are also clearly seen, as are the structures due to the overlay of the asymmetric stretches due to the loss of CO_2 and gain of CO_2^* , see Fig. S10 which shows the spectrum collected at $100\text{ }^\circ\text{C}$ in Fig. 6a, and Fig. S11 which compares the bands obtained in the thermal and plasma experiments directly. As expected, the P and R bands due to CO are absent from the thermal spectra and, given the noise on the spectra in Fig. 6a–c is ca. 4×10^{-4} , this suggests the CO absorption at 2116 cm^{-1} must be $\leq 4 \times 10^{-4}$, hence $[\text{CO}] \leq 1.2 \times 10^{-5}\text{ M}$ and the conversion of CO_2 to CO is $\leq 0.1\%$, i.e. very low, as would be expected on the basis of the literature [41] and in complete contrast to the plasma experiments. The bands around 2900 cm^{-1} are due to the protective polymer film on the beam splitter of the spectrometer.

Conclusions

The combination of the FTIR reflectance and transmission plasma cells allowed reasonable estimations of the conversion of CO_2 to CO, and provided a direct method of assessing the temperature of the gases in the plasma glow using the broadening of the CO absorption. The conversion and energy efficiencies observed were comparable to those routinely reported in the literature, giving confidence in the approach. Further, and for the first time, the same chemical system was studied driven both by plasma and thermally, and the data compared and contrasted: whilst up to 9% conversion of the CO_2 to CO was observed in the plasma experiments, no conversion was observed in the thermal experiment up to $600\text{ }^\circ\text{C}$. Finally, strong plasma-induced absorptions were observed on the Ti/Macor which remained for some time after the plasma was stopped which remain unassigned, but have been tentatively attributed to transverse and longitudinal optical modes of SiO_2 and work is continuing to obtain definitive identification of these and the implications of their formation.

In summary, we believe the experimental methodology described has significant potential to elucidate the chemistry taking place in non-thermal plasma catalysis and hence provide the paradigm shift required to take the technology forward to commercial reality.

Acknowledgements Abd Halim Bin Md Ali would like to thank the Government of Malaysia for a scholarship to study for a Ph.D., the staff of the Majlis Amanah Rakyat in Kuala Lumpur, Malaysia and London, United Kingdom for their invaluable assistance. Z. T. A. W. Mashhadani would like to thank the Government of Iraq for a scholarship, the staff of the Higher Committee for Education Development in Iraq for their invaluable assistance and the Iraq Ministry of Oil Petroleum Research and Development Center PRDC for the opportunity to study for a Ph.D. The authors also wish to thank Mr. Neville Dickman and Mr. S. Daley for their assistance and support for this work.

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